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# 第一章 Introduction



# 第二章 Demos

## 2.1 Chemical formulae

$\text{H}_2\text{O}$ ,  $\text{Sb}_2\text{O}_3$

$\text{H}^+$ ,  $\text{CrO}_4^{2-}$ ,  $[\text{AgCl}_2]^-$ ,  $\text{Y}^99+$ ,  $\text{Y}^{99+}$

$\text{Fe}^{\text{II}}\text{Fe}^{\text{III}}\text{O}_4$ ,  $2\text{H}_2\text{O}$ ,  $2\text{H}_2\text{O}$ ,  $0.5\text{H}_2\text{O}$ ,  $\frac{1}{2}\text{H}_2\text{O}$ ,  $(_{1/2})\text{H}_2\text{O}$ ,  $n\text{H}_2\text{O}$

$^{227}_{90}\text{Th}^+$ ,  $^{227}_{90}\text{Th}^+$ ,  $^{01}_{-1}\text{n}^-$ ,  $^{01}_{-1}\text{n}^-$

$\text{H}^3\text{HO}$ ,  $\text{H}^3\text{HO}$

$(\text{NH}_4)_2\text{S}$ ,  $[(\text{X}_2)_3]^{3+}$

$\text{H}_2(\text{aq})$ ,  $\text{CO}_3^{2-}(\text{aq})$ ,  $\text{NaOH}(\text{aq}, \infty)$

$\text{OCO}^{\cdot-}$ ,  $\text{NO}^{(2\cdot)-}$

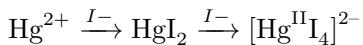
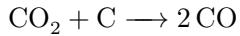
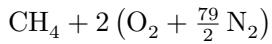
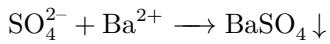
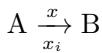
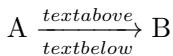
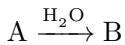
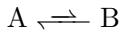
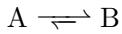
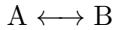
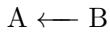
$\text{NO}_x$ ,  $\text{NO}_x$ ,  $\text{Fe}^{n+}$ ,  $\text{Fe}^{n+}$

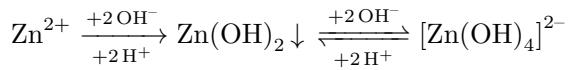
$\text{Pt}(\text{C}_2\text{H}_4)\text{Cl}_3^-$

$\text{KCr}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}$ ,  $\text{KCr}(\text{SO}_4)_{2.12}\text{H}_2\text{O}$ ,  $\text{KCr}(\text{SO}_4)_2 * 12\text{H}_2\text{O}$

$\text{C}_6\text{H}_5\text{CHO}$ ,  $\text{A}-\text{B}=\text{C}\equiv\text{D}$ ,  $\text{A}\equiv\text{B}\equiv\text{C}\equiv\text{D}$ ,  $\text{A}\cdots\text{B}\cdots\text{C}$ ,  $\text{A}\rightarrow\text{B}\leftarrow\text{C}$

## 2.2 Chemical equations





$$K = \frac{[\text{Hg}^{2+}][\text{Hg}]}{[\text{Hg}_2^{2+}]}$$

$$K = \frac{[\text{Hg}^{2+}][\text{Hg}]}{[\text{Hg}^{2+}]}$$

## 2.3 Structural formulae

