

title

author

2018-09-22

Contents

1	Introduction	5
2	Demos	7
2.1	Chemical formulae	7
2.2	Chemical equations	7
2.3	Structural formulae	8

Chapter 1

Introduction

Chapter 2

Demos

2.1 Chemical formulae

H_2O , Sb_2O_3

H^+ , CrO_4^{2-} , $[\text{AgCl}_2]^-$, Y^{99+} , Y^{99+}

$\text{Fe}^{\text{II}}\text{Fe}^{\text{III}}\text{O}_4$, $2\text{H}_2\text{O}$, $2\text{H}_2\text{O}$, $0.5\text{H}_2\text{O}$, $\frac{1}{2}\text{H}_2\text{O}$, $(_{1/2})\text{H}_2\text{O}$, $n\text{H}_2\text{O}$

$^{227}_{90}\text{Th}^+$, $^{227}_{90}\text{Th}^+$, $^{01}_{-1}\text{n}^-$, $^{01}_{-1}\text{n}^-$

H^3HO , H^3HO

$(\text{NH}_4)_2\text{S}$, $[(\text{X}_2)_3]_2^{3+}$

$\text{H}_2(\text{aq})$, $\text{CO}_3^{2-}(\text{aq})$, $\text{NaOH}(\text{aq}, \infty)$

$\text{OCO}^{\cdot -}$, $\text{NO}^{(2\cdot)-}$

NO_x , Fe^{n+} , $\text{Fe}^{\text{n}+}$,
 $\text{Pt}(\text{C}_2\text{H}_4)\text{Cl}_3^-$

$\text{KCr}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}$, $\text{KCr}(\text{SO}_4)_{2.12}\text{H}_2\text{O}$, $\text{KCr}(\text{SO}_4)_2 * 12\text{H}_2\text{O}$

$\text{C}_6\text{H}_5-\text{CHO}$, $\text{A}-\text{B}=\text{C}\equiv\text{D}$, $\text{A}-\text{B}=\text{C}\# \text{D}$

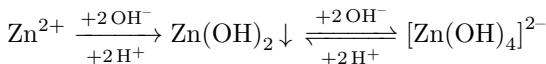
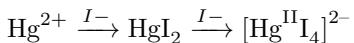
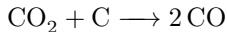
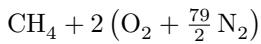
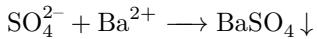
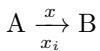
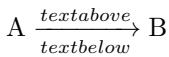
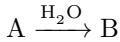
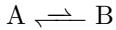
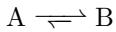
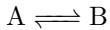
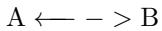
$\text{A}-\text{B}=\text{C}\equiv\text{D}$, $\text{A}=\text{B}=\text{C}\equiv\text{D}$, $\text{A}\cdots\text{B}\cdots\text{C}$, $\text{A}\rightarrow\text{B}\leftarrow\text{C}$

2.2 Chemical equations

$\text{A} \longrightarrow \text{B}$

$\text{A} \longleftarrow \text{B}$

$\text{A} \longleftrightarrow \text{B}$



$$K = \frac{[Hg^{2+}][Hg]}{[Hg_2^{2+}]}$$

$$K = \frac{[Hg^{2+}][Hg]}{[Hg_2^{2+}]}$$

2.3 Structural formulae

